

GCE

Chemistry A

H432/03: Unified chemistry

A Level

Mark Scheme for June 2022

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It is also responsible for developing new specifications to meet national requirements and the needs of students and teachers. OCR is a not-for-profit organisation; any surplus made is invested back into the establishment to help towards the development of qualifications and support, which keep pace with the changing needs of today's society.

This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which marks were awarded by examiners. It does not indicate the details of the discussions which took place at an examiners' meeting before marking commenced.

All examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes should be read in conjunction with the published question papers and the report on the examination.

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MARKING INSTRUCTIONS**PREPARATION FOR MARKING****RM ASSESSOR**

1. Make sure that you have accessed and completed the relevant training packages for on-screen marking: *RM Assessor Online Training*; *OCR Essential Guide to Marking*.
2. Make sure that you have read and understood the mark scheme and the question paper for this unit.
3. Log-in to RM Assessor and mark the **required number** of practice responses (“scripts”) and the **required number** of standardisation responses.

MARKING

1. Mark strictly to the mark scheme.
2. Marks awarded must relate directly to the marking criteria.
3. The schedule of dates is very important. It is essential that you meet the RM Assessor 50% and 100% (traditional 50% Batch 1 and 100% Batch 2) deadlines. If you experience problems, you must contact your Team Leader (Supervisor) without delay.
4. If you are in any doubt about applying the mark scheme, consult your Team Leader by telephone, email or via the RM Assessor messaging system.
5. Work crossed out:

Crossed Out Responses

Where a candidate has crossed out a response and provided a clear alternative then the crossed out response is not marked. Where no alternative response has been provided, examiners may give candidates the benefit of the doubt and mark the crossed out response where legible.

Rubric Error Responses – Optional Questions

Where candidates have a choice of question across a whole paper or a whole section and have provided more answers than required, then all responses are marked and the highest mark allowable within the rubric is given. Enter a mark for each question answered into RM assessor, which will select the highest mark from those awarded. *(The underlying assumption is that the candidate has penalised themselves by attempting more questions than necessary in the time allowed.)*

Multiple Choice Question Responses

When a multiple choice question has only a single, correct response and a candidate provides two responses (even if one of these responses is correct), then no mark should be awarded (as it is not possible to determine which was the first response selected by the candidate).

When a question requires candidates to select more than one option/multiple options, then local marking arrangements need to ensure consistency of approach.

Contradictory Responses

When a candidate provides contradictory responses, then no mark should be awarded, even if one of the answers is correct.

Short Answer Questions (requiring only a list by way of a response, usually worth only **one mark per response**)

Where candidates are required to provide a set number of short answer responses then only the set number of responses should be marked. The response space should be marked from left to right on each line and then line by line until the required number of responses have been considered. The remaining responses should not then be marked. Examiners will have to apply judgement as to whether a 'second response' on a line is a development of the 'first response', rather than a separate, discrete response. *(The underlying assumption is that the candidate is attempting to hedge their bets and therefore getting undue benefit rather than engaging with the question and giving the most relevant/correct responses.)*

Short Answer Questions (requiring a more developed response, worth **two or more marks**)

If the candidates are required to provide a description of, say, three items or factors and four items or factors are provided, then mark on a similar basis – that is downwards (as it is unlikely in this situation that a candidate will provide more than one response in each section of the response space.)

Longer Answer Questions (requiring a developed response)

Where candidates have provided two (or more) responses to a medium or high tariff question which only required a single (developed) response and not crossed out the first response, then only the first response should be marked. Examiners will need to apply professional judgement as to whether the second (or a subsequent) response is a 'new start' or simply a poorly expressed continuation of the first response.

6. Always check the pages (and additional objects if present) at the end of the response in case any answers have been continued there. If the candidate has continued an answer there then add a tick to confirm that the work has been seen.

7. Award No Response (NR) if:

- there is nothing written in the answer space.

Award Zero '0' if:

- anything is written in the answer space and is not worthy of credit (this includes text and symbols).

Team Leaders must confirm the correct use of the NR button with their markers before live marking commences and should check this when reviewing scripts.

8. The RM Assessor **comments box** is used by your Team Leader to explain the marking of the practice responses. Please refer to these comments when checking your practice responses. **Do not use the comments box for any other reason.**

If you have any questions or comments for your Team Leader, use the phone, the RM Assessor messaging system, or email.

9. Assistant Examiners will send a brief report on the performance of candidates to their Team Leader (Supervisor) via email by the end of the marking period. The report should contain notes on particular strengths displayed as well as common errors or weaknesses. Constructive criticism of the question paper/mark scheme is also appreciated.

10. For answers marked by levels of response:

Read through the whole answer from start to finish, using the Level descriptors to help you decide whether it is a strong or weak answer. The indicative scientific content in the Guidance column indicates the expected parameters for candidates' answers, but be prepared to recognise and credit unexpected approaches where they show relevance. Using a 'best-fit' approach based on the skills and science content evidenced within the answer, first decide which set of level descriptors, Level 1, Level 2 or Level 3, best describes the overall quality of the answer.

Once the level is located, award the higher or lower mark:

The higher mark should be awarded where the level descriptor has been evidenced and all aspects of the communication statement (in italics) have been met.

The lower mark should be awarded where the level descriptor has been evidenced but aspects of the communication statement (in italics) are missing.

In summary:

The skills and science content determines the level.

The communication statement determines the mark within a level.

Level of response questions on this paper are **3a** and **5a**.

The only annotation on a level of response question should be the indication of the level.

A level annotation should be used where all marks for a level have been achieved e.g. a candidate has 6 marks, so they would have this annotation on their script:

L3

If a candidate has achieved 5 marks then they have reached Level 3 but with one mark omitted. They should have the following annotations on their scripts:

L3 **^**















The same principle should be applied to Level 2 and Level 1.

No marks (0) should have a cross: **×**

Place the annotations alongside the mark for the question.

On additional pages, annotate using **SEEN**

11. Annotations available in RM Assessor

Annotation	Meaning
	Correct response
	Incorrect response
	Omission mark
	Benefit of doubt given
	Contradiction
	Rounding error
	Error in number of significant figures
	Error carried forward
	Level 1
	Level 2
	Level 3
	Benefit of doubt not given
	Noted but no credit given
	Ignore

12. Abbreviations, annotations and conventions used in the detailed Mark Scheme (to include abbreviations and subject-specific conventions).

Annotation	Meaning
DO NOT ALLOW	Answers which are not worthy of credit
IGNORE	Statements which are irrelevant
ALLOW	Answers that can be accepted
()	Words which are not essential to gain credit
—	Underlined words must be present in answer to score a mark
ECF	Error carried forward
AW	Alternative wording
ORA	Or reverse argument

13. Subject-specific Marking Instructions

INTRODUCTION

Your first task as an Examiner is to become thoroughly familiar with the material on which the examination depends. This material includes:

- the specification, especially the assessment objectives
- the question paper
- the mark scheme.

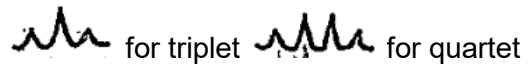

You should ensure that you have copies of these materials.

You should ensure also that you are familiar with the administrative procedures related to the marking process. These are set out in the OCR booklet **Instructions for Examiners**. If you are examining for the first time, please read carefully **Appendix 5 Introduction to Script Marking: Notes for New Examiners**.

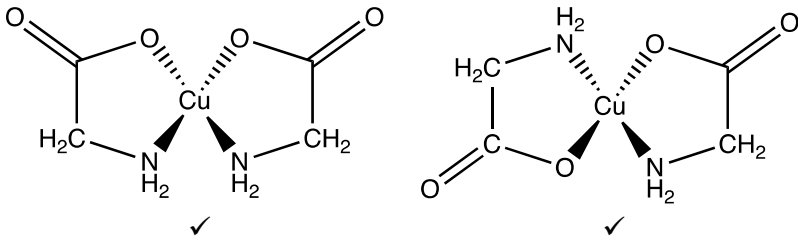
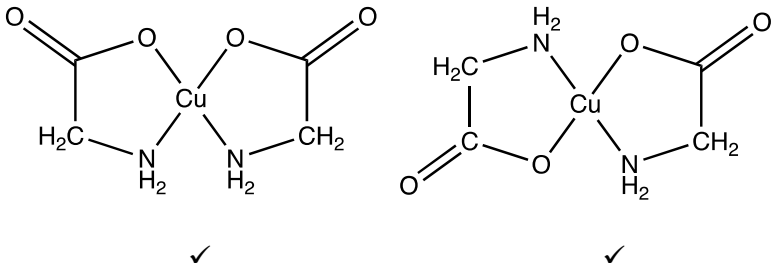

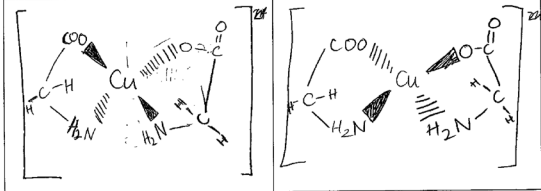
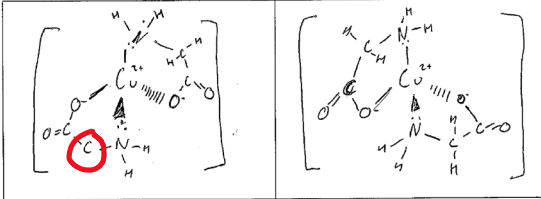
Please ask for help or guidance whenever you need it. Your first point of contact is your Team Leader.

Question			Answer	Marks	AO element	Guidance
1	(a)	(i)	<p>Structure and bonding NH_3 is (simple) molecular/simple covalent/ /has intermolecular forces AND NH_4NO_3 is ionic ✓</p> <p>Comparison of strength Ionic bonds are stronger than intermolecular bonds / forces between molecules OR Ionic bonds need more energy to break than intermolecular bonds ✓</p>	2	AO1.1 ×2	<p>For intermolecular bonds/forces ALLOW hydrogen bonds OR London Forces/induced dipole forces/permanent dipole forces OR van der Waals' forces</p> <p>ALLOW NH_4NO_3 has molecular ions NH_4^+ and NO_3^- are molecular ions</p> <p>ORA</p> <p>ALLOW: Intermolecular bonds are weak AND ionic bonds are strong ✓</p>
		(ii)	<p>(NH_4^+) nitrogen has oxidation number of -3 AND (NO_3^-) nitrogen has oxidation number of +5 ✓</p> <p><i>i.e. nitrogens are -3 AND +5 gets the mark BOTH signs essential</i></p>	1	AO1.2	<p>Statement that one student is correct is NOT required. <i>Implicit in answer</i></p> <p>ALLOW 3- AND 5+</p>

Question		Answer	Marks	AO element	Guidance	
	(b)	(i)	Cu: 66% AND Zn 34% ✓	1	AO2.6	
		(ii)	<p>FIRST CHECK ANSWER ON THE ANSWER LINE If answer = 65.42 (to 2 DP) award 2 marks</p> <p>-----</p> <p>Numerator from Zn isotopes $(64 \times 16.82) + (66 \times 9.53) + (67 \times 1.38) + (68 \times 6.27)$ OR 2224.28 ✓</p> <p>Relative atomic mass Numerator \div 34 AND answer to 2 DP ✓ Mark ECF from numerator</p> <p>$\frac{(64 \times 16.82) + (66 \times 9.53) + (67 \times 1.38) + (68 \times 6.27)}{34} \checkmark$</p> <p>= 65.42 (to 2 DP) ✓</p>	2	AO1.2 ×2	<p>Refer to answer to 1b(i) for ECF from incorrect % composition of Zn and Cu</p> <p>ECF \div by Zn % in b(i)</p> <p>-----</p> <p>Common errors</p> <p>22.24 \div100 and answer to 2 DP \rightarrow 1 mark for numerator</p> <p>64.23 All 6 isotopes used \rightarrow No marks</p> <p>188.91 All 6 isotopes used \rightarrow 6423 for numerator \div34 and 2 DP \rightarrow 1 mark by ECF</p>

Question		Answer	Marks	AO element	Guidance															
(c)	(i)	<table border="1"> <thead> <tr> <th>Proton environment</th> <th>Splitting pattern</th> <th></th> </tr> </thead> <tbody> <tr> <td>1</td> <td>Triplet</td> <td>Triplet AND quartet ✓</td> </tr> <tr> <td>2</td> <td>Quartet</td> <td></td> </tr> <tr> <td>3</td> <td>Doublet</td> <td>Doublet AND triplet ✓</td> </tr> <tr> <td>4</td> <td>Triplet</td> <td></td> </tr> </tbody> </table>	Proton environment	Splitting pattern		1	Triplet	Triplet AND quartet ✓	2	Quartet		3	Doublet	Doublet AND triplet ✓	4	Triplet		2	AO1.2 ×2	<p>For quartet, ALLOW Quad.... e.g. quadruplet, quadlet, quadret, etc</p> <p>For doublet, ALLOW duplet</p> <p>ALLOW diagrams to show splitting pattern e.g.</p> <p> for triplet  for quartet</p> <p>ALLOW splitting patterns shown as numbers i.e. '3' for triplet, '4' for quartet</p>
Proton environment	Splitting pattern																			
1	Triplet	Triplet AND quartet ✓																		
2	Quartet																			
3	Doublet	Doublet AND triplet ✓																		
4	Triplet																			
(c)	(ii)	<p>Environment 2: (Protons) adjacent to (one) C=O ✓</p> <p>Environment 3: (Protons) adjacent/between/surrounded by 2 C=O / a ketone AND aldehyde OR C=O on both sides ✓</p>	2	AO3.1 ×2	<p>ALLOW HC–C=O</p> <p>DO NOT ALLOW H–C=O</p> <p>DO NOT ALLOW HC–O <i>Simply reading $\delta = 3.6$ ppm from data sheet)</i></p> <p>IGNORE 'next to 2 Os'</p>															

Question		Answer	Marks	AO element	Guidance
	(d) (i)	<p>Bond angles $\text{H}_2\text{NCH}_2\text{COONa}$, bond angle = 107° AND $\text{HOOCCH}_2\text{NH}_3\text{Cl}$, bond angle = 109.5° ✓</p> <p>Number of electron pairs Mark independently of angles</p> <p>In $\text{NaOH}/107^\circ$, (NH_2 has) 3 bonded pairs / 3 bonds AND 1 lone pair ✓</p> <p>In $\text{HCl}/109.5^\circ$, (NH_3^+ has) 4 bonded pairs / 4 bonds ✓</p>	3	AO1.2 ×3	<p>ALLOW 107 ± 0.5</p> <p>ALLOW 109 OR 110°</p> <p>ALLOW NH_2 has 4 pairs, one of which is a lone pair</p> <p>For bonded pairs/bonds ALLOW bonded groups, atoms, elements, regions Bonded essential</p> <p>IGNORE electron region OR electron density</p> <p>IGNORE NH_3 has no lone pairs</p> <p>IGNORE lone pairs repel more (than bonded pairs)</p> <p>IGNORE shapes, even if wrong</p> <p>ALLOW bp for bonded pair and lp for lone pair</p>
	(ii)	<p>Equation: $2 \text{H}_2\text{NCH}_2\text{COOH} + \text{Cu}(\text{CH}_3\text{COO})_2$ $\rightarrow \text{Cu}(\text{H}_2\text{NCH}_2\text{COO})_2 + 2 \text{CH}_3\text{COOH}$ ✓</p>	3	AO2.6	<p>ALLOW molecular formulae or mixture, e.g. $2\text{C}_2\text{H}_5\text{NO}_2 + \text{CuC}_4\text{H}_6\text{O}_4$ $\rightarrow \text{CuC}_4\text{H}_8\text{N}_2\text{O}_4 + 2\text{C}_2\text{H}_4\text{O}_2$</p> <p>IGNORE charges</p>

Question	Answer	Marks	AO element	Guidance
	<p>Structures</p>  <p>OR</p>  <p>Ligands must be shown as bidentate rings IGNORE connectivity for NH₂ BUT connectivity must be to O of COO</p>		<p>AO2.5 ×2</p>	<p>i.e. IGNORE wrong or missing charges in ionic compounds if formula is correct/ e.g. ALLOW Cu(CH₃COO⁻)₂, Cu⁺(CH₃COO⁻)₂</p> <p>ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous</p> <p>IGNORE charges</p> <p>ALLOW arc to represent -CH₂- between: C of C=O and NH₂</p>  <p>ALLOW 1 mark for 2 'correct' structures shown as tetrahedral e.g.</p>  <p>IGNORE missing Hs on C, e.g.</p> 

Question			Answer	Marks	AO element	Guidance
2	(a)	(i)	<p>FIRST CHECK ANSWER ON THE ANSWER LINE If answer = 2.19×10^{-3} award 3 marks</p> <p>-----</p> <p>$n(\text{Cl}_2) = 420/24 = 17.5 \text{ (mol)} \checkmark$</p> <p>$n(\text{Ca}(\text{C/O})_2) = \frac{17.5}{2} = 8.75 \text{ (mol)} \checkmark$</p> <p>Concentration $\text{Ca}(\text{C/O})_2 = \frac{8.75}{4 \times 1000}$ $= 2.19 \times 10^{-3} \text{ (mol dm}^{-3}\text{)} \checkmark$ 3SF AND standard form</p>	3	AO2.2 ×3	<p>Use of ideal gas equation for all 3 marks provided 'sensible' p and T used: e.g. from 101 kPa and 298 K $\rightarrow n = 17.122 \rightarrow 2.14 \times 10^{-3}$ from 100 kPa and 298 K $\rightarrow n = 16.952 \rightarrow 2.12 \times 10^{-3}$ Examples of 'sensible' $p = 100 \text{ kPa, } 101 \text{ kPa, } 101,325 \text{ Pa}$ $T = 273 - 298 \text{ K}$</p> <p>ALLOW ECF</p> <p>-----</p> <p>Common errors</p> <p>$4.38 \times 10^{-3} \text{ (no } \div 2) \rightarrow 2 \text{ marks}$</p> <p>$2.19 \times 10^n \rightarrow 2 \text{ marks}$</p> <p>$4.38 \times 10^n \rightarrow 1 \text{ mark}$</p> <p>$2.2 \times 10^{-3} \rightarrow 2 \text{ marks}$ <i>not appropriate SF</i></p>

Question	Answer	Marks	AO element	Guidance
	<p>(ii) Equation $3 \text{Ca(ClO)}_2 \rightarrow 2 \text{CaCl}_2 + \text{Ca(ClO}_3)_2 \checkmark$</p> <p>Reduction Cl reduced from +1 to -1 ✓</p> <p>Oxidation Cl oxidised from +1 to +5 ✓</p> <p>+1 starting oxidation number seen once Cl required for both explanation marks</p> <p>IGNORE oxidation numbers shown below/above equation <i>(treat as rough working)</i> BUT If no oxidation numbers in explanation, <i>look at equation for oxidation numbers</i></p>	3	AO2.6 AO1.2 ×2	<p>ALLOW multiples ALLOW $3 \text{ClO}^- \rightarrow 2 \text{Cl}^- + \text{ClO}_3^-$</p> <p>ALLOW 1 out of 2 redox marks if oxidation number changes are BOTH correct ...BUT reduction/oxidation is incorrectly assigned, i.e. Cl is oxidised from +1 to -1 Cl is reduced from +1 to +5</p> <p>ALLOW 1 out of 2 redox marks if oxidation changes correct but red and ox not stated Cl changes from +1 to -1 Cl changes from +1 to +5</p> <p>----- General: ALLOW number before sign in ox no, e.g. 1- for -1</p> <p>IGNORE ionic charges, e.g. Cl^{5+} IGNORE '1' (signs required)</p> <p>IGNORE references to electron loss/gain (even if wrong)</p>

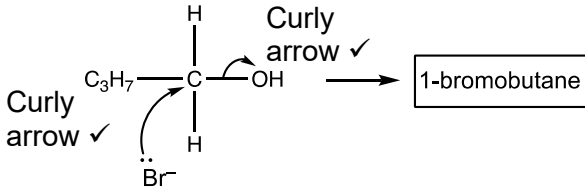
Question	Answer	Marks	AO element	Guidance
(b)	<p>6 marking points → 5 MAX</p> <p>-----</p> <p>ALLOW labels 1, 2 and 3; A, B and C, etc, provided that meaning is clear</p> <p>-----</p> <p>Oxidising agent AND equation</p> <p>$\text{Cr}_2\text{O}_7^{2-}$ is oxidising agent with $\text{C}_2\text{H}_5\text{OH}$ /oxidises $\text{C}_2\text{H}_5\text{OH}$ ✓</p> <p>$\text{Cr}_2\text{O}_7^{2-} + 8\text{H}^+ + 3\text{C}_2\text{H}_5\text{OH} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O} + 3\text{CH}_3\text{CHO}$ ✓</p> <p>Explanation for $\text{Cr}_2\text{O}_7^{2-}/\text{Cr}^{3+}$ and $\text{CH}_3\text{CHO}/\text{C}_2\text{H}_5\text{OH}$</p> <p>$E$ for $\text{Cr}_2\text{O}_7^{2-}/\text{Cr}^{3+}$ is more +ve /higher /greater</p> <p>OR</p> <p>$E_{\text{cell}} = (+)1.527 \text{ V}$ + sign not required</p> <p>OR</p> <p>$\text{Cr}_2\text{O}_7^{2-}/\text{Cr}^{3+}$ equilibrium shifts right ✓</p> <p>Reducing agent AND equation</p> <p>Cr^{3+} is reducing agent with FeO_4^{2-} /reduces FeO_4^{2-} ✓</p> <p>$2\text{Cr}^{3+} + 2\text{H}^+ + 2\text{FeO}_4^{2-} \rightarrow \text{Cr}_2\text{O}_7^{2-} + \text{H}_2\text{O} + 2\text{Fe}^{3+}$ ✓</p> <p>Explanation for $\text{Cr}_2\text{O}_7^{2-}/\text{Cr}^{3+}$ and $\text{FeO}_4^{2-}/\text{Fe}^{3+}$</p> <p>$E$ for $\text{Cr}_2\text{O}_7^{2-}/\text{Cr}^{3+}$ is less +ve (E) / lower /smaller</p> <p>OR</p> <p>$E_{\text{cell}} = (+)0.87 \text{ V}$ + sign not required</p> <p>OR</p> <p>$\text{Cr}_2\text{O}_7^{2-}/\text{Cr}^{3+}$ equilibrium shifts left ✓</p>	5	<p>AO2.5</p> <p>AO2.6</p> <p>AO2.6</p> <p>AO2.5</p> <p>AO2.6</p> <p>AO2.6</p>	<p>ALLOW reverse argument (ORA) throughout</p> <p>For equations, ALLOW multiples</p> <p>In equations, ALLOW \rightleftharpoons for \rightarrow</p> <p>ALLOW $\text{Cr}_2\text{O}_7^{2-}$ is oxidising agent if linked to $\text{C}_2\text{H}_5\text{OH}$ as reactant in equation</p> <p>ALLOW Cr^{6+} for $\text{Cr}_2\text{O}_7^{2-}$</p> <p>ALLOW $\text{Cr}_2\text{O}_7^{2-}$ is reduced by $\text{C}_2\text{H}_5\text{OH}$</p> <p>In explanation, look for CONS between 'OR' statements</p> <p>ALLOW Cr^{3+} is reducing agent if clearly linked to FeO_4^{2-} as reactant in equation</p> <p>ALLOW Fe^{6+} for FeO_4^{2-}</p> <p>ALLOW Cr^{3+} is oxidised by FeO_4^{2-}</p> <p>In explanation, look for CONS between 'OR' statements</p> <p>-----</p>

Question		Answer	Marks	AO element	Guidance
					<p>Note on equations There are 2 marks for the equations with H⁺, H₂O and e⁻ cancelled down</p> <p>ALLOW 1 mark for 2 'correct' equations where H⁺, H₂O and e⁻ have NOT all been cancelled down.</p> <p><i>e.g. 1 mark from 2 uncanceled equations</i> $\text{Cr}_2\text{O}_7^{2-} + 14\text{H}^+ + 3\text{C}_2\text{H}_5\text{OH}$ $\rightarrow 2\text{Cr}^{3+} + 6\text{H}^+ + 7\text{H}_2\text{O} + 3\text{CH}_3\text{CHO}$</p> $2\text{Cr}^{3+} + 2\text{H}^+ + 2\text{FeO}_4^{2-} + 6\text{e}^-$ $\rightarrow \text{Cr}_2\text{O}_7^{2-} + \text{H}_2\text{O} + 2\text{Fe}^{3+} + 6\text{e}^-$
	(c)	$5 \text{H}_2\text{S} + 2 \text{MnO}_4^- + 6 \text{H}^+ \rightarrow 2 \text{Mn}^{2+} + 5 \text{S} + 8 \text{H}_2\text{O}$ <p>OR</p> $40 \text{H}_2\text{S} + 16 \text{MnO}_4^- + 48 \text{H}^+ \rightarrow 16 \text{Mn}^{2+} + 5 \text{S}_8 + 64 \text{H}_2\text{O}$ <p>Any FIVE correct species ✓</p> <p>Correct balanced equation ✓</p>	2	AO3.2	<p>ALLOW multiples e.g. $2\frac{1}{2} \text{H}_2\text{S} + \text{MnO}_4^- + 3 \text{H}^+$ $\rightarrow \text{Mn}^{2+} + 2\frac{1}{2} \text{S} + 4 \text{H}_2\text{O}$</p> $20 \text{H}_2\text{S} + 8 \text{MnO}_4^- + 24 \text{H}^+$ $\rightarrow 8 \text{Mn}^{2+} + 2\frac{1}{2} \text{S}_8 + 32 \text{H}_2\text{O}$ <p>IGNORE extra species containing: Mn, H, S and O ONLY BUT ALLOW KMnO₄ on LHS, forming K⁺ on RHS</p> <p>IGNORE electrons</p> <p>IGNORE state symbols</p>

Question	Answer	Marks	AO element	Guidance
3 (a)*	<p><i>Refer to marking instructions on page 4 of mark scheme for guidance on marking this question.</i></p> <p>Level 3 (5-6 marks) Diagram showing reflux with most labels AND A CORRECT calculation of the % yield of 1-bromobutane AND A detailed description of most purification steps.</p> <p><i>There is a well-developed line of reasoning which is clear and logically structured. The information presented is relevant and substantiated.</i></p> <p>Level 2 (3-4 marks) Diagram showing reflux with some labels AND Calculates the % yield of 1-bromobutane with some errors OR Diagram showing reflux with most labels AND describes some purification steps, with some detail OR Calculates the % yield of 1-bromobutane with some errors AND describes some purification steps, with some detail</p> <p><i>There is a line of reasoning presented with some structure. The information presented is relevant and supported by some evidence.</i></p>	6	AO2.8 ×2 AO3.3 ×4	<p>Indicative scientific points may include:</p> <p>Diagram Diagram draw with condenser above flask Labels including</p> <ul style="list-style-type: none"> condenser water in at bottom and out at top pear-shaped or round-bottom flask <p>Calculation of % yield of 1-bromobutane</p> <ul style="list-style-type: none"> $n(\text{butan-1-ol}) = \frac{9.25}{74.0} = 0.125 \text{ (mol)}$ mass 1-bromobutane = $6.10 \times 1.268 = 7.7348 \text{ g}$ $n(1\text{-bromobutane}) = \frac{7.7348}{136.9} = 0.0565 \text{ (mol)}$ $\% \text{ yield} = \frac{0.0565}{0.125} \times 100 = 45.2\%$ <p>ALLOW 45.2 ± 0.2 for small slip/rounding -----</p> <p>NOTE Use of 6.1 g (omission of density)</p> <ul style="list-style-type: none"> $n(1\text{-bromobutane}) = \frac{6.10}{136.9} = 0.044558... \text{ (mol)}$ $\% \text{ yield} = \frac{0.044558...}{0.125} \times 100 = 35.6\%$ <p>Purification</p> <ul style="list-style-type: none"> In separating funnel, organic layer is on bottom Drying with an anhydrous salt by formula or name, e.g. MgSO₄, Na₂SO₄, CaCl₂ Redistil at 102°C <p>Examples of detail in bold (NOT INCLUSIVE)</p> <p>NOTE: 'Use a separating funnel', dry, and 'redistil' on their own are NOT detailed descriptions</p>

Question	Answer	Marks	AO element	Guidance
	<p>Level 1 (1-2 marks) Diagram showing reflux OR Attempts to calculate the % yield of 1-bromobutane OR Describes few purification steps.</p> <p><i>There is an attempt at a logical structure with a line of reasoning. The information is in the most part relevant.</i></p> <p>0 marks No response or no response worthy of credit.</p>			

Question	Answer	Marks	AO element	Guidance
(b)	<p>Step 1 The oxygen atom of the alcohol group accepts a proton to form a positively-charged intermediate. 2 marks</p> <p>Step 2 Bromide ions react with the intermediate by nucleophilic substitution to form 1-bromobutane. 2 marks</p> <p>2 possible routes:</p> <p>EITHER</p> <p>OR</p>	4	AO3.2 x4	<p>ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous</p> <p>For CH₃CH₂CH₂, ALLOW CH₃(CH₂)₂, C₃H₇</p> <p>IGNORE dipoles</p> <p>-----</p> <p>ALLOW curly arrow to H of H-O-SO₃H OR H-Br</p> <p>IGNORE absence of curly arrow from H-O or from H-Br</p> <p>+ charge MUST be on O of intermediate</p> <p>Curly arrow must</p> <ul style="list-style-type: none"> start from, OR be traced back to any point across width of lone pair on :Br⁻ OR :OH OR start from - charge on Br⁻ <p>(Lone pair NOT needed if curly arrow shown from - charge on Br⁻)</p> <p>IGNORE final products: 1-bromobutane and H₂O</p> <p>IF C₃H₇CH₂-O⁺H₂ is not shown, ALLOW intermediate mark for carbocation: C₃H₇CH₂⁺</p>

Question	Answer	Marks	AO element	Guidance
				<p>ALLOW 2 marks max for mechanism without positively charge intermediate, i.e.</p>  <p>If in doubt, contact Team Leader</p>

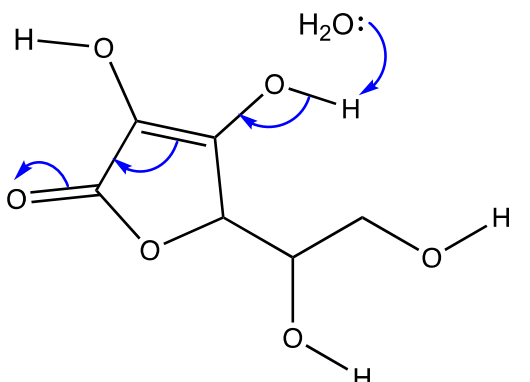
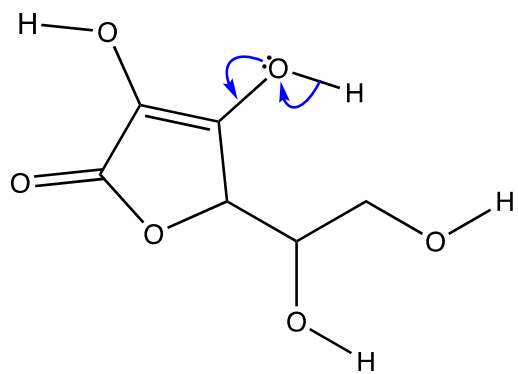
Question		Answer	Marks	AO element	Guidance
4	(a)	<p>At 90 °C/higher temperature</p> <ul style="list-style-type: none"> Faster rate AND more frequent collisions ✓ More particles have the activation energy/E_a or greater ✓ $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ is lower ✓ (forward reaction) ΔH –ve OR exothermic ✓ 	4	<p>AO2.7 ×1</p> <p>AO1.2 ×1</p> <p>AO2.3 ×1</p> <p>AO1.2 ×1</p>	<p>ORA for 50 °C</p> <p>IGNORE more successful collisions</p> <p>ALLOW more molecules have enough energy to react</p> <p>ALLOW atoms/molecules/ions</p> <p>ALLOW decreases</p>
	(b)	(i)			
		<p>Cl^- /It/They react with AgNO_3 / Ag^+ /silver ions</p> <p>OR</p> <p>AgCl formed</p> <p>OR</p> <p>$\text{Ag}^+ + \text{Cl}^- \rightarrow \text{AgCl}$ ✓</p>	1	AO3.2	<p>IGNORE chlorine/Cl for chloride ion</p> <p>IGNORE AgCl_2</p>

Question		Answer	Marks	AO element	Guidance
	(ii)	<p>$[\text{CoCl}_4^{2-}]$ decreases AND $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ increases ✓</p> <p>C_l^- increase is $4 \times$ change in $[\text{CoCl}_4^{2-}] / [\text{Co}(\text{H}_2\text{O})_6]^{2+}$ ✓</p> <p>Equilibrium shifts to right ✓</p>	3	<p>AO3.1 $\times 2$</p> <p>AO3.2 $\times 1$</p>	<p>IGNORE missing charges and small slips in formulae, e.g. CoCl_4 missing bracket, etc</p> <p>IGNORE C_l^- for changes in concentration</p> <p>ALLOW suitable alternatives for 'shifts to right', e.g. towards products OR in forward direction OR 'favours the right'</p>

Question	Answer	Marks	AO element	Guidance
5 (a)*	<p><i>Please refer to the marking instructions on page 4 of this mark scheme for guidance on how to mark this question.</i></p> <p>Level 3 (5–6 marks) Calculates correct enthalpy change with correct – sign for $\Delta_{\text{hy}}H(\text{Ca}^{2+})$, allowing for acceptable errors.</p> <p><i>There is a well-developed line of reasoning which is clear and logically structured. The information presented is relevant and substantiated.</i></p> <p>Level 2 (3–4 marks) Calculates a value of $\Delta_{\text{sol}}H(\text{CaCl}_2(\text{s}))$ from the: Energy change AND Amount in mol of CaCl_2.</p> <p><i>There is a line of reasoning presented with some structure. The information presented is relevant and supported by some evidence.</i></p> <p>Level 1 (1–2 marks) Processes experimental data to obtain the: Energy change from $mc\Delta T$ OR Amount in mol of CaCl_2.</p> <p><i>There is an attempt at a logical structure with a line of reasoning. The information is in the most part relevant.</i></p> <p>0 marks – No response or no response worthy of credit.</p>	6	AO3.1 ×4 AO3.2 ×2	<p><i>Indicative scientific points may include:</i></p> <p>1. Processing experimental data Energy change from $mc\Delta T$</p> <ul style="list-style-type: none"> Energy in J OR kJ = $106.6 \times 4.18 \times 18.5 = 8243.378$ (J) OR 8.243378 (kJ) <p><i>3SF or more</i></p> <p>Amount in mol of CaCl_2</p> <ul style="list-style-type: none"> $n(\text{CaCl}_2) = \frac{9.28}{111.1} = 0.0835\dots\dots$ (mol) <p>0.08352835284 unrounded</p> <p>-----</p> <p>2. \pm value of $\Delta_{\text{sol}}H(\text{CaCl}_2(\text{s}))$</p> <p>= $\pm \frac{8.24\dots}{0.0835\dots} = \pm 98.68957929$ (kJ mol⁻¹)</p> <p><i>3 SF or more. From 3 SF: $\frac{8.24}{0.0835} = 98.7$</i></p> <p>-----</p> <p>3. CORRECT $\Delta_{\text{hy}}H(\text{Ca}^{2+})$ calculated with signs</p> <p>$\Delta_{\text{hy}}H(\text{Ca}^{2+}) = \text{L.E.} + \Delta_{\text{sol}}H(\text{CaCl}_2) - 2 \Delta_{\text{hy}}H(\text{Cl}^-)$ = $-2223 + (-98.7) - (2 \times -378)$ = -1566 (kJ mol⁻¹)</p> <p><i>3SF or more with correct – sign</i></p> <p>From unrounded values, -1565.689579</p> <p>-----</p> <p>See next page for examples of acceptable errors</p>

Question			Answer	Marks	AO element	Guidance
						<p><u>Acceptable errors</u></p> <p>ALLOW omission of trailing zeroes</p> <p>ALLOW minor slips in rounding, transcription errors, etc throughout</p> <p>ALLOW one small error, e.g. subtracting mass of CaCl₂ for <i>m</i> $m = 106.60 - 9.28 = 97.32$ $q = 7.5257556 \text{ (kJ)}$ $\Delta_{\text{sol}}H = 90.09821629 \text{ (kJ mol}^{-1}\text{)}$ $\Delta_{\text{hy}}H(\text{Ca}^{2+}) = -1557 \text{ (kJ mol}^{-1}\text{)}$</p> <p>OR adding mass of CaCl₂ for <i>m</i> $m = 106.60 + 9.28 = 115.88$ $q = 8.9610004 \text{ kJ}$ $\Delta_{\text{sol}}H = 107.2809423 \text{ (kJ mol}^{-1}\text{)}$ $\Delta_{\text{hy}}H(\text{Ca}^{2+}) = -1574 \text{ (kJ mol}^{-1}\text{)}$</p>
	(b)	(i)	$\text{C}_8\text{H}_{18} + \text{C}_2\text{H}_5\text{OH} + 15\frac{1}{2} \text{O}_2 \rightarrow 10 \text{CO}_2 + 12 \text{H}_2\text{O} \checkmark$	1	AO2.6	<p>ALLOW multiples e.g. $2 \text{C}_8\text{H}_{18} + 2 \text{C}_2\text{H}_5\text{OH} + 31 \text{O}_2$ $\rightarrow 20 \text{CO}_2 + 24 \text{H}_2\text{O}$</p> <p>ALLOW C₁₀H₂₄O for C₈H₁₈ + C₂H₅OH <i>Combining ethanol and octane!</i></p>

Question	Answer	Marks	AO element	Guidance
	<p>(ii) FIRST CHECK ANSWER ON THE ANSWER LINE If answer = 341850 to 2 SF or more award 3 marks</p> <p>-----</p> <p>$M(\text{C}_8\text{H}_{18}) = 114$ AND $M(\text{C}_2\text{H}_5\text{OH}) = 46$ OR 1 mol C_8H_{18} + 1 mol $\text{C}_2\text{H}_5\text{OH}$ has mass of 160 g ✓</p> <p>50 mol C_8H_{18} OR 50 mol $\text{C}_2\text{H}_5\text{OH}$ OR 50 mol $(\text{C}_8\text{H}_{18} + \text{C}_2\text{H}_5\text{OH})$ OR 8.00 kg fuel contains 50 mol C_8H_{18} + 50 mol $\text{C}_2\text{H}_5\text{OH}$ ✓</p> <p>Energy = $(50 \times 5470) + (50 \times 1367)$ OR $50 \times (5470 + 1367)$ OR 50×6837 OR $273500 + 68350$</p> <p>= 341850 (kJ) ✓</p>	3	AO2.2 ×3	<p>IGNORE sign throughout</p> <p>ALLOW approach based on mass for 2nd mark</p> <p>$m(\text{C}_8\text{H}_{18}) = (114/160) \times 8000 = 5700$ g AND $m(\text{C}_2\text{H}_5\text{OH}) = (46/160) \times 8000 = 2300$ g</p> <p>Energy = $5700/114 \times 5470 + 2300/46 \times 1367$ = 341850 (kJ)</p> <p>ALLOW 2 SF or more correctly rounded</p> <p>-----</p> <p>Common errors</p> <p>310800 → 2 marks Use of equal masses (4 kg) of C_8H_{18} & $\text{C}_2\text{H}_5\text{OH}$ (rather than equal moles)</p> <p>Example</p> <p>energy released when 4kg of C_8H_{18} burnt</p> <p>$\frac{4000}{114} = 35 \text{ moles} \dots$ $35 \times 5470 = 191425 \text{ KJ released}$</p> <p>$\frac{4000}{46} = 87 \text{ moles} \dots$ $87 \times 1367 = 118869 \text{ KJ}$ 191425 $191425 + 118869 = 310294$ energy released = ... 310300 kJ [3]</p>

Question		Answer	Marks	AO element	Guidance
6	(a)	Number of optical isomers = 4 ✓	1	AO2.1	
	(b) (i)	Hydrogen bonding AND Many OH/hydroxyl / hydroxy / alcohol ✓	1	AO2.1	ALLOW 4 OH DO NOT ALLOW OH ⁻
	(ii)	x = 15 ✓ y = 31 ✓	2	AO3.2 ×2	
	(c) (i)	 <p>3 OR 4 curly arrows correct → 2 marks ✓✓ 1 curly arrow correct → 1 mark ✓</p>	2	AO3.2 ×2	<p>IGNORE incorrect curly arrows</p> <p>IGNORE 'double' curly arrows such as:</p>  <p>H₂O Curly arrow must</p> <ul style="list-style-type: none"> start from, OR be traced back to any point across width of lone pair on H₂O:

Question	Answer	Marks	AO element	Guidance
(ii)	<p>FIRST CHECK ANSWER ON THE ANSWER LINE If answer = 2.16 award 3 marks</p> <p>-----</p> <p>[Vitamin C] = $0.150 \times 4 = 0.600$ (mol dm⁻³) ✓ <i>0.6 seen anywhere</i></p> <p>[H⁺] = $\sqrt{K_a \times [\text{Vitamin C}]}$ = $\sqrt{7.94 \times 10^{-5} \times 0.600}$ = 6.90×10^{-3} (mol dm⁻³) ✓</p> <p>pH = $-\log [\text{H}^+]$ = $-\log 6.90 \times 10^{-3}$ = 2.16 ✓ 2 DP required</p>	3	<p>AO2.4 ×2</p> <p>AO1.2 ×1</p>	<p>For [H⁺] ALLOW ECF from incorrect [vitamin C]</p> <p>for pH ALLOW ECF ONLY if [H⁺] has been derived from K_a AND [vitamin C]</p> <p>-----</p> <p>COMMON ERRORS</p> <p>pH = 4.32 2/3 calculation marks <i>No square root of $(7.94 \times 10^{-5} \times 0.600)$</i></p> <p>pH = 2.46 2/3 calculation marks <i>No $\times 4$ $(7.94 \times 10^{-5} \times 0.150)$</i></p> <p>pH = 2.76 2/3 calculation marks <i>$\div 4$ $(7.94 \times 10^{-5} \times 0.0375)$</i></p> <p>pH = 4.92 1/3 calculation mark <i>No square root AND 0.150</i></p> <p>pH = 5.53 1/3 calculation mark <i>No square root AND 0.0375</i></p>

Question	Answer	Marks	AO element	Guidance
(d)	<p>FIRST CHECK ANSWER ON THE ANSWER LINE If answer = 38 (mg) award 4 marks</p> <p>-----</p> $n(I_2) = 22.50 \times \frac{9.60 \times 10^{-4}}{1000} = 2.16 \times 10^{-5} \text{ (mol) } \checkmark$ <p>$n(\text{vitamin C})$ in 250 cm³ volumetric flask $= 10 \times 2.16 \times 10^{-5} = 2.16 \times 10^{-4} \text{ (mol) } \checkmark$</p> <p>$M(\text{Vitamin C: C}_6\text{H}_8\text{O}_6) = 176 \text{ OR } (12 \times 6) + (1 \times 8) + (16 \times 6)$ <i>Seen anywhere</i></p> <p>Mass vitamin C in 150 cm³ of orange $= 2.16 \times 10^{-4} \times 176.0 = 0.038016 \text{ g}$ $= 38 \text{ (mg) } \checkmark$ 2 SF or more</p>	4	AO2.8 ×4	<p>Use ECF throughout Intermediate values for working to at least 3 SF. TAKE CARE as value written down may be truncated value stored in calculator. Depending on rounding, either can be credited.</p> <p>-----</p> <p>COMMON ERRORS:</p> <p>22.81 mg scaling by 150/250 → 3 marks FINAL MARK LOST BY SCALING</p> <p>Determine the mass, in mg, of vitamin C in a 150 cm³ serving of the orange juice.</p> $0.0225 \times 9.60 \times 10^{-4} = 2.16 \times 10^{-5} \text{ mol } \checkmark \quad \text{C}_6\text{H}_8\text{O}_6$ $2.16 \times 10^{-5} \times 10 = 2.16 \times 10^{-4} \text{ mol } \checkmark \quad = 176 \checkmark$ $\frac{2.16 \times 10^{-4}}{0.250} = 8.64 \times 10^{-4} \text{ mol dm}^{-3}$ $8.64 \times 10^{-4} \times 0.15 = 1.296 \times 10^{-4} \text{ mol}$ $1.296 \times 10^{-4} \times 176 = 0.0228096 \text{ g}$ $0.0228 \times 1000 = 22.81 \text{ } \times$ <p>42.24 mg using 25.0 cm³ instead of 22.50 → 3 marks</p> <p>25.34 mg using 25.0 cm³ AND scaling by 150/250 instead of 22.50 → 2 marks</p> <p>63.36 mg scaling by 250/150 → 3 marks</p>

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